

Download Ph And Poh Chart Answer Key

Part 2: For each of the problems below, assume 100% dissociation. 1. A. Write the equation for the dissociation of hydrochloric acid. B. Find the pH of a 0.00476 M hydrochloric acid solution. 2. A. pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral ... pOH = $-\log [\text{OH}^-]$ So if $[\text{OH}^-] = 10^{-8}$, pOH = 8 Together, $\text{pH} + \text{pOH} = 14$. Complete the following chart. [H⁺] pH [OH⁻] pOH Acidic or Basic

1. 10^{-5} M	5	10^{-9} M	9	Acidic
2. 10^{-7} M	7	10^{-7} M	7	Neutral
3. 10^{-10} M	10	10^{-4} M	4	Basic
4. 10^{-2} M	2			

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